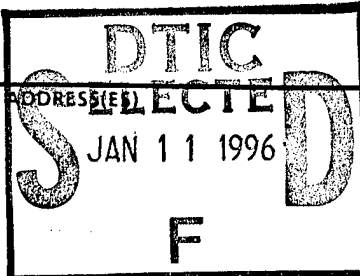


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**SYNTHESIS AND STRUCTURAL CHARACTERIZATION OF ALKYL
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OF $(\text{Me}_3\text{CCH}_2)_2(\text{Cl})\text{Ga}\cdot\text{P}(\text{SiMe}_3)_3$, $\text{R}_2\text{GaP}(\text{SiMe}_3)_2\text{GaR}_2\text{Cl}$ ($\text{R} = \text{Me}_3\text{CCH}_2$ and
 Me_3SiCH_2), and $[(\text{R})(\text{X})\text{GaP}(\text{SiMe}_3)_2]_2$ ($\text{R} = \text{Me}_3\text{CCH}_2$, $\text{X} = \text{Cl}$; $\text{R} = \text{Me}_3\text{CCH}_2$,
 $\text{X} = \text{Me}_3\text{CCH}_2$; $\text{R} = \text{Me}_3\text{SiCH}_2$, $\text{X} = \text{Br}$)**

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Abstract- Reaction of $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ with $(\text{Me}_3\text{Si})_3\text{P}$ in a 1:1 mole ratio does not result in any intermolecular dehalosilylation but affords the adduct $(\text{Me}_3\text{CCH}_2)_2(\text{Cl})\text{Ga}\cdot\text{P}(\text{SiMe}_3)_3$ (**1**) in a nearly quantitative yield. Mixing $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ and $(\text{Me}_3\text{Si})_3\text{P}$ in a 2:1 mole ratio yields as the only isolable product the dimeric compound, $[(\text{Me}_3\text{CCH}_2)(\text{Cl})\text{GaP}(\text{SiMe}_3)_2]_2$ (**2**), containing three different substituents on the gallium center. This type of product is also observed in the 2:1 mole reaction of $(\text{Me}_3\text{SiCH}_2)_2\text{GaBr}$ and $(\text{Me}_3\text{Si})_3\text{P}$, where $[(\text{Me}_3\text{SiCH}_2)(\text{Br})\text{GaP}(\text{SiMe}_3)_2]_2$ (**3**) is obtained. The salt elimination reaction between $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ and $\text{LiP}(\text{SiMe}_3)_2$ leads to the dimer $[(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2]_2$ (**4**) in a good yield. Equilibration of compound **4** with 2 molar equivalents of $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ produces the mixed-bridge $(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{CMe}_3)_2\text{Cl}$ (**5**) in a nearly quantitative yield. Mixed-bridge compound $(\text{Me}_3\text{SiCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{SiMe}_3)_2\text{Cl}$ (**6**) results from the 2:1 mole ratio dehalosilylation reaction of $(\text{Me}_3\text{SiCH}_2)_2\text{GaCl}$ and $(\text{Me}_3\text{Si})_3\text{P}$. Various physical and spectroscopic data are presented for compounds **1** - **6**, as well as the results of their X-ray crystal structure determinations.

Introduction

Continued activity in the development of single-source precursors to 13-15 semiconductor materials¹ has motivated our laboratory to investigate the synthesis of novel organogallium-phosphorus compounds which might serve as potential precursors to GaP. Recently, our efforts to produce the gallium-phosphorus bond have led to the formation of interesting ring compounds and simple adducts. For example, the ring compounds $[\text{X}_2\text{GaP}(\text{SiMe}_3)_2]_2$ ($\text{X} = \text{Cl}^2, \text{Br}^3, \text{I}^3$), and $\text{Ph}_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{Ph})_2\text{Cl}$ (**7**)⁴ have been isolated in our laboratory through dehalosilylation reactions, along with the adducts $\text{R}_3\text{Ga}\cdot\text{P}(\text{SiMe}_3)_3$ [$\text{R} = \text{Ph}$ (**8**)⁴ and Me_3SiCH_2 (**9**)⁵] and $\text{Ph}_2(\text{Cl})\text{Ga}\cdot\text{P}(\text{SiMe}_3)_3$ (**10**)⁴. As part of a continuing effort to expand our knowledge of the gallium-phosphorus system, we wanted to further investigate reactions involving $(\text{Me}_3\text{Si})_3\text{P}$ and its monolithium salt $\text{LiP}(\text{SiMe}_3)_2$ with alkylgallium halides. The bulky neopentyl and (trimethylsilyl)methyl ligands were chosen as the alkyl substituents on the gallium atom in hopes of affecting the oligomerization or structural conformation of the resulting products. Our goal was to produce compounds which could serve as effective intermediates to species containing novel

mixtures of Group 13 and Group 15 atoms. Through these efforts, we were able to isolate several new compounds and to observe some interesting chemical phenomena; herein, we report the results obtained from our investigations.

Experimental Section

General Considerations. All manipulations of air- and moisture-sensitive materials were performed in a Vacuum Atmospheres HE-493 Dri-Lab containing an argon atmosphere or by general Schlenk techniques. Benzene was distilled from sodium benzophenone ketyl under dry dinitrogen. Pentane was distilled over LiAlH_4 under dry dinitrogen. $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ ⁶, $(\text{Me}_3\text{SiCH}_2)_2\text{GaCl}$ ⁷, $(\text{Me}_3\text{SiCH}_2)_2\text{GaBr}$ ⁷ and $(\text{Me}_3\text{Si})_3\text{P}$ ⁸ were synthesized by literature procedures. $\text{LiP}(\text{SiMe}_3)_2$ was prepared via the 1:1 mole reaction of $(\text{Me}_3\text{Si})_3\text{P}$ and MeLi . The integrity of all materials used was confirmed via ^1H NMR spectra and by melting points. ^1H , $^{13}\text{C}\{^1\text{H}\}$, and $^{31}\text{P}\{^1\text{H}\}$ NMR spectra were recorded on a Varian XL-300 spectrometer operating at 300, 75.4, and 121.4 MHz, respectively. ^1H and $^{13}\text{C}\{^1\text{H}\}$ spectra were referenced to TMS by using the residual protons or carbons of deuterated benzene at δ 7.15 or 128 ppm, respectively, or the upfield pentet of C_7D_8 at δ 2.09 ppm for ^1H NMR spectra. All NMR samples were prepared in 5-mm tubes which were septum sealed under argon. Melting points (uncorrected) were obtained with a Thomas-Hoover Uni-melt apparatus, and capillaries were flame-sealed under argon. Elemental Analyses were performed by E+R Microanalytical Laboratory, Inc., Corona, NY. Mass spectral data were collected on a JEOL JMS-SX 102A spectrometer operating in the electron ionization mode at 20 eV. X-ray crystallographic data for compounds **1** - **6** were obtained as follows: **1** - **3** obtained at -135 °C on a Rigaku AFC6/S diffractometer utilizing graphite-monochromated Mo-K α ($\lambda = 0.71073$ Å) radiation; **4** - **5** obtained at 25 °C on a Siemens P4 diffractometer utilizing graphite-monochromated Mo-K α ($\lambda = 0.71073$ Å) radiation; **6** obtained at -140 °C on a Rigaku AFC-7R diffractometer using graphite-monochromated Mo-K α ($\lambda = 0.71073$ Å) radiation

Preparation of $(\text{Me}_3\text{CCH}_2)_2(\text{Cl})\text{Ga}\cdot\text{P}(\text{SiMe}_3)_3$ (1**).** $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ (0.247 g, 1.00 mmol) in 25 mL of pentane and $(\text{Me}_3\text{Si})_3\text{P}$ (0.250 g, 0.998 mmol) in 25 mL of pentane were

combined in a 250 mL one-necked round-bottomed flask equipped with a Teflon valve and a magnetic stirbar. The resulting clear solution was allowed to stir at room temperature outside of the drybox for 48 h. The volatiles were removed *in vacuo* to yield a white crystalline solid which was dissolved in 5 mL of warm pentane. Cooling the resulting solution to -15 °C for several days afforded colorless crystals of **1** suitable for X-ray analysis (0.487 g, 98% yield); mp. 89-96 °C. Anal. Calcd. (found) for $C_{19}H_{49}ClGaPSi_3$: C, 45.82 (46.78); H, 9.92 (9.67); Cl, 7.12 (6.96). 1H NMR (C_6D_6): δ 0.29 [d, $SiMe_3$ ($J_{P-H} = 4.8$ Hz)], 1.14 (s, CH_2), 1.32 (s, $C-Me_3$). $^{13}C\{^1H\}$ NMR (C_6D_6): δ 3.33 [d, $SiMe_3$ ($J_{P-C} = 8.4$ Hz)], 33.03 (s, $-CMe_3$), 34.38 (s, Me_3), 40.39 (s, CH_2). $^{31}P\{^1H\}$ NMR: δ -230.12. Mass spec.: m/z 496 ($C_{19}H_{47}ClGaPSi_3$) $^+ = (M - 2H)^+$

Preparation of $[(Me_3CCH_2)(Cl)GaP(SiMe_3)_2]_2$ (**2**).

[Compound **2** was synthesized using a procedure similar to that used for **1**]

Reactants: $(Me_3CCH_2)_2GaCl$ (0.495g, 2.00 mmol), $(Me_3Si)_3P$ (0.250g, 0.998 mmol).

Product: 0.163 g, 46% yield (based on P), mp. 29-44 °C. X-ray quality crystals were grown from a pentane/chlorobenzene mixture. Anal. Calcd. (found) for $C_{22}H_{58}Cl_2Ga_2P_2Si_4$: C, 37.30 (36.83); H, 8.26 (8.19). 1H NMR (C_6D_6): δ 0.50 [t, $P-SiMe_3$ ($J_{P-H} = 3.6$ Hz)], 1.21 (s, CH_2), 1.30 (s, CMe_3). $^{13}C\{^1H\}$ NMR (C_6D_6): δ 3.21 (s, CMe_3), 4.13 (s, CH_2), 6.77 [t, $P-SiMe_3$ ($J_{P-C} = 4.7$ Hz)]. $^{31}P\{^1H\}$ NMR: δ -233.86.

Preparation of $[(Me_3SiCH_2)(Br)GaP(SiMe_3)_2]_2$ (**3**).

[Compound **3** was synthesized using a procedure similar to that used for **1**]

Reactants: $(Me_3SiCH_2)_2GaBr$ (0.648g, 2.00 mmol), $(Me_3Si)_3P$ (0.250g, 0.998 mmol).

Product: 0.170 g, 41% yield (based on P), mp. 87-101 °C. X-ray quality crystals were grown from pentane. Anal. Calcd. (found) for $C_{20}H_{58}Br_2Ga_2P_2Si_6$: C, 28.99 (29.18); H, 7.06 (7.11). 1H NMR (C_6D_6): δ 0.24 (s, CH_2), 0.35 (s, $SiMe_3$), 0.50 [t, $P-SiMe_3$ ($J_{P-H} = 2.7$ Hz)]. $^{13}C\{^1H\}$ NMR (C_6D_6): δ 2.42 (s, $SiMe_3$), 2.80 (s, CH_2), 3.87 [t, $P-SiMe_3$ ($J_{P-C} = 3.7$ Hz)]. $^{31}P\{^1H\}$ NMR: δ -227.61. Mass spec.: m/z 829 ($C_{20}H_{59}Br_2Ga_2P_2Si_6$) $^+ = (M + H)^+$

Preparation of $[(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2]_2$ (4). $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ (0.247 g, 1.00 mmol) in 25 mL of benzene and $\text{LiP}(\text{SiMe}_3)_2$ (0.184 g, 0.999 mmol) in 25 mL of benzene were combined in a 250 mL one-necked round-bottomed flask equipped with a Teflon valve and a magnetic stirbar. The resulting yellow solution was allowed to stir at room temperature outside of the drybox for 48 h. The volatiles were removed *in vacuo* to yield a yellow powder which was extracted with 25 mL of pentane. Upon filtration, concentration and cooling of the extract to -15°C for several days, colorless crystals of **4** suitable for X-ray analysis were obtained (0.618 g, 79% yield); mp. $49 - 58^\circ\text{C}$. Anal. Calcd. (found) for $\text{C}_{32}\text{H}_{80}\text{Ga}_2\text{P}_2\text{Si}_4$: C, 49.94 (50.11); H, 10.33 (10.68). ^1H NMR (C_6D_6): δ 0.42 [t, SiMe_3 ($J_{\text{P-H}} = 4.8$ Hz)], 1.11 (s, CH_2), 1.25 (s, C-Me $_3$). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 3.89 [t, SiMe_3 ($J_{\text{P-C}} = 2.6$ Hz)], 32.82 [d, Me $_3$ ($J_{\text{P-C}} = 3.8$ Hz)], 34.82 (s, -CMe $_3$). $^{31}\text{P}\{^1\text{H}\}$ NMR: δ -215.24.

Preparation of $(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{CMe}_3)_2\text{Cl}$ (5).

Compound **4** (0.637 g, 1.00 mmol) in 25 mL of pentane and $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ (0.594 g, 2.00 mmol) in 25 mL of pentane were combined in a 250 mL one-necked round-bottomed flask equipped with a Teflon valve and a magnetic stirbar. The resulting clear solution was allowed to stir at room temperature outside of the drybox for 48 h. The volatiles were removed *in vacuo* to yield a white crystalline solid which was dissolved in 5 mL of warm pentane. Cooling the resulting solution to -30°C for several days afforded colorless crystals of **5** suitable for X-ray analysis (1.206 g, 98% yield); mp. $44-48^\circ\text{C}$. Anal. Calcd. (found) for $\text{C}_{26}\text{H}_{62}\text{ClGa}_2\text{PSi}_2$: C, 49.04 (48.84); H, 9.81 (9.69); Cl, 5.57 (5.25). ^1H NMR (C_6D_6): δ 0.38 [d, SiMe_3 ($J_{\text{P-H}} = 4.8$ Hz)], 1.24 (s, C-Me $_3$), 1.30 (s, CH_2). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 4.59 [d, SiMe_3 ($J_{\text{P-C}} = 7.6$ Hz)], 34.55 (s, Me $_3$), 39.67 [d, CH_2 ($J_{\text{P-C}} = 14.0$ Hz)]. $^{31}\text{P}\{^1\text{H}\}$ NMR: δ -210.22.

Preparation of $(\text{Me}_3\text{SiCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{SiMe}_3)_2\text{Cl}$ (6).

[Compound **6** was synthesized using a procedure similar to that used for **1**]

Reactants: $(\text{Me}_3\text{SiCH}_2)_2\text{GaCl}$ (0.559 g, 2.00 mmol), $(\text{Me}_3\text{Si})_3\text{P}$ (0.250 g, 0.998 mmol).

Product: 0.694 g, 75% yield, mp. $68-75^\circ\text{C}$. X-ray quality crystals were grown from pentane.

Anal. Calcd. (found) for $\text{C}_{22}\text{H}_{62}\text{ClGa}_2\text{PSi}_6$: C, 37.68 (37.48); H, 8.84 (8.74); P, 4.57 (4.18).

^1H NMR (C_6D_6): δ 0.25 (s, $-\text{CH}_2\text{SiMe}_3$), 0.34 [d, SiMe_3 ($J_{\text{P-H}} = 5.4$ Hz)]. $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 2.32 (s, $-\text{SiMe}_3$), 3.37 [d, P-SiMe_3 ($J_{\text{P-C}} = 7.8$ Hz)], 7.81 [d, CH_2 ($J_{\text{P-C}} = 10.7$ Hz)]. $^{31}\text{P}\{^1\text{H}\}$ NMR: δ -213.27. Mass spec.: m/z 697 ($\text{C}_{22}\text{H}_{61}\text{ClGa}_2\text{PSi}_6$) $^+ = (\text{M} - \text{H})^+$

X-ray structural solution and refinement

Crystallographic data are summarized in Table 1. The structural analyses were performed as follows: Compounds **1** - **3**: Colorless crystals of **1**, **2**, and **3** were mounted on a glass fiber with a viscous oil under a stream of cold dinitrogen. X-ray intensity data were recorded at -135 °C on a Rigaku AFC6/S diffractometer utilizing graphite-monochromated Mo- $\text{K}\alpha$ radiation ($\lambda = 0.71073$ Å), and the structures were solved by direct methods. Full-matrix least-squares refinement with weights based upon counting-statistics was performed. Hydrogen atoms were incorporated at their calculated positions using a riding model in the later iterations of refinement which converged at $R = 0.043$ ($R_w = 0.051$) for **1**, $R = 0.063$ ($R_w = 0.075$) for **2**, and $R = 0.036$ ($R_w = 0.041$) for **3**. A final difference-Fourier synthesis revealed no unusual features. Crystallographic calculations were performed using the NRCVAX⁹ suite of structure determination programs. For all structure-factor calculations, neutral atom scattering factors and their anomalous dispersion corrections were taken from ref. 10. Interatomic distances and angles are given in Tables 2 - 3. ORTEP¹¹ diagrams showing the solid-state conformations and atom numbering schemes of **1**, **2**, and **3** are presented in Figures 1 - 3. Compounds **4** - **5**: Colorless crystals of **4** and **5** suitable for X-ray diffraction were mounted in glass capillaries under argon. The unit-cell parameters were obtained by the least-squares refinement of the angular settings of 24 reflections ($20^\circ < 2\theta < 25^\circ$). The structures were solved by direct methods, completed by subsequent difference Fourier syntheses and refined by full-matrix least-squares procedures. All non-hydrogen atoms were refined anisotropically. Hydrogen atoms were treated as idealized contributions. All software and sources of the scattering factors are contained in the SHELXTL PLUS¹² (4.2) and SHELXTL (5.3) program libraries. ORTEP diagrams showing the solid-state conformations and atom numbering schemes of **4** and **5** are presented in Figures 4 - 5. Compound **6**: A colorless parallelepiped crystal of **6** was mounted in a glass capillary under argon atmosphere. Cell parameters and an orientation matrix for data collection, obtained from a least-squares analysis of the setting angles of 24 reflections in the range

$42.48^\circ < 2\Theta < 49.21^\circ$, corresponded to a monoclinic cell. The structure was solved by direct methods. The non-hydrogen atoms were refined anisotropically. Hydrogen atoms were located by standard difference Fourier techniques and were included in the structure factor calculation at idealized positions based on a riding model. An isotropic group thermal parameter was refined for all of the hydrogens. Structure solution, refinement, and the calculation of derived results were performed using the TEXSAN¹³ and SHELXTL¹² package of computer programs. Neutral atom scattering factors were those of Cromer and Waber¹⁴, and the real and imaginary anomalous dispersion corrections were those of Cromer¹⁵. An ORTEP diagram showing the solid-state conformation and atom numbering scheme of **6** is presented in Figure 6.

Results and Discussion

Tris(trimethylsilyl)phosphine, $(\text{Me}_3\text{Si})_3\text{P}$, reacts at room temperature with $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ in a 1:1 mole ratio to form the simple Lewis acid-base adduct $(\text{Me}_3\text{CCH}_2)_2(\text{Cl})\text{Ga}\cdot\text{P}(\text{SiMe}_3)_3$ (**1**) in a nearly quantitative yield. This result is in accord with that seen when the alkyl/aryl group on the gallium center is phenyl⁴. Compound **1** was characterized by ^1H , $^{13}\text{C}\{^1\text{H}\}$, and $^{31}\text{P}\{^1\text{H}\}$ NMR spectroscopy, partial elemental analysis, and single-crystal X-ray analysis (vide infra).

The 2:1 mole ratio reaction of $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ and $(\text{Me}_3\text{Si})_3\text{P}$ does not undergo intermolecular dehalosilylation to afford the desired mixed-bridge compound containing a four-membered $\overline{\text{Ga-P-Ga-Cl}}$ ring, but yields instead $[(\text{Me}_3\text{CCH}_2)(\text{Cl})\text{GaP}(\text{SiMe}_3)_2]_2$ (**2**) as the only isolable product. This is in contrast to the results observed when the substituent on the gallium atom is phenyl, where the mixed-bridge species is formed in a moderate yield⁴. Isolation of **2** as the sole product of the reaction is surprising yet plausible. Ligand redistributions of this type are not without precedent in 13-15 systems and can be observed most notably in the aluminum-phosphorus system¹⁶.

Similarly, $[(\text{Me}_3\text{SiCH}_2)(\text{Br})\text{GaP}(\text{SiMe}_3)_2]_2$ (**3**) was the sole product obtained from the 2:1 mole reaction of $(\text{Me}_3\text{SiCH}_2)_2\text{GaBr}$ with $(\text{Me}_3\text{Si})_3\text{P}$, and was recovered in a 41% crystalline yield. This compound was characterized by ^1H , $^{13}\text{C}\{^1\text{H}\}$, and $^{31}\text{P}\{^1\text{H}\}$ NMR, partial elemental analysis, electron impact mass spectrometry, and a single-crystal X-ray analysis (vide infra). The

formation pathway of **3** is believed to be similar to that of **2**. It is important to note that compounds **2** and **3** are unique in that they are gallium compounds which contain three different substituents on the metal center. As evidenced by Beachley and co-workers¹⁷, these compounds are difficult to fully characterize in the solid and solution phases. To this end, dimers **2** and **3** represent only the third and fourth examples, in the Ga-P system, of these types of compounds to be characterized by X-ray crystallographic means. Those whose solid-state structure have been determined are $\{(\text{tBu})[\text{P}(\text{H})(\text{tBu}_3\text{C}_6\text{H}_2)]\text{GaCl}\}_2$ (**11**)¹⁸ and $[(\text{Me}_3\text{CCH}_2)(\text{Cl})\text{GaPPh}_2]_3$ (**12**)¹⁷.

Dimer $[(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2]_2$ (**4**) is obtained from the direct 1:1 mole ratio salt-elimination reaction of $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ and $\text{LiP}(\text{SiMe}_3)_2$ in benzene at room temperature. The solution ^1H NMR data are consistent with the solid-state structure of **4** as revealed by the single-crystal X-ray analysis. Most notably, the spectrum contains a triplet centered at δ 0.42 ppm which corresponds to virtual coupling between the ring phosphorus atoms and the SiMe_3 protons, indicating the dimeric nature of the compound in solution. The absence of a doublet resonance in the ^1H NMR spectrum suggests that **4** does not dissociate in solution to give any monomeric species. Dimeric compounds in the gallium-phosphorus system which contain two alkyl or aryl groups on the metal centers and two trimethylsilyl groups on the pnictogen atoms are scarce; and, to our knowledge, compound **4** represents only the second example of a structurally characterized species of this type [the other being $[\text{Me}_2\text{GaP}(\text{SiMe}_3)_2]_2$ (**13**)].

Equilibration of dimer **4** with 2 molar equivalents of $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ leads to the mixed-bridge $(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{CMe}_3)_2\text{Cl}$ (**5**). As noted above, compound **5** could not be isolated from the direct 2:1 mole ratio reaction of $(\text{Me}_3\text{CCH}_2)_2\text{GaCl}$ and $(\text{Me}_3\text{Si})_3\text{P}$. Dimer to mixed-bridge interconversions have been noted in several 13-15 systems²⁰. With this in mind, we had hoped to utilize this synthetic method to arrive at our desired mixed-bridge.

In contrast to the above, the mixed-bridge species $(\text{Me}_3\text{SiCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{SiMe}_3)_2\text{Cl}$ (**6**) is obtained from the direct 2:1 mole ratio reaction of $(\text{Me}_3\text{SiCH}_2)_2\text{GaCl}$ with $(\text{Me}_3\text{Si})_3\text{P}$ in pentane at room temperature. This compound was characterized by a variety of methods including partial elemental analysis, electron impact mass spectrometry, and single-crystal X-ray analysis. The ^1H , $^{13}\text{C}\{^1\text{H}\}$, and $^{31}\text{P}\{^1\text{H}\}$ NMR spectra are consistent with those expected for compounds such as **6**. Most notably, in the ^1H

NMR, a doublet resonance at δ 0.34 corresponding to coupling of the ^{31}P nucleus with the methyl protons of the adjacent trimethylsilyl groups and a singlet resonance at δ 0.25 corresponding to the methylene and methyl protons of the (trimethylsilyl)methyl are observed. The EI mass spectrum shows a fragmentation pattern at $m/z = 697$, which corresponds to the molecular ion of **6** minus a proton. Interestingly, halogen mixed-bridge compounds in the gallium phosphorus system are rare, and **5** and **6** represent only the second and third examples of these species to be characterized in the solution and solid phase. It is important to note that attempted 1:1 mole ratio reactions of $(\text{Me}_3\text{SiCH}_2)_2\text{GaCl}$ with $(\text{Me}_3\text{Si})_3\text{P}$ and $\text{LiP}(\text{SiMe}_3)_2$ did not produce any characterizable compounds, and only the starting materials were recovered in each case.

Crystallographic data and data collection parameters for **1** - **6** are summarized in Table 1. ORTEP diagrams showing the atom-numbering schemes for **1** - **6** are presented in Figures 1 -6, respectively. Tables 2 - 7 list selected bond distances and bond angles, while tables 8 - 9 list selected non-hydrogen atom fractional coordinates and equivalent isotropic thermal parameters for **1** - **3** and **4** - **6**, respectively.

Adduct **1** crystallizes in the monoclinic system with four molecules occupying the general positions of the centrosymmetric space group $\text{P}2_1/\text{c}$. In the solid state, this molecule adopts the staggered conformation depicted in Figure 1. The Ga-P bond distance at 2.534(2) Å in **1** is slightly longer than that of 2.459(2) Å found in **10**. This is not surprising due to the larger steric bulk of the neopentyl groups in **1** to that of the phenyl ligands in **10**. This trend is also evidenced in the Ga-Cl and mean Ga-C bond lengths [Ga-Cl = 2.258 Å in **1** and 2.240 Å in **10**; Ga-C (mean) = 1.999 Å in **1** and 1.982 Å in **10**]. The Cl-Ga-P angle in **1** at 97.49(6)° is much smaller than that of 101.1(1)° seen in **10**, as is the mean P-Ga-C angle [103.8° in **1**, 111.2° in **10**].

Crystals of dimer **2** belong to the monoclinic system, space group $\text{P}2_1/\text{c}$, with the P atoms lying on a crystallographic C_2 axis of symmetry, and thus the Ga-P-Ga'-P' ring is strictly planar. The mean Ga-P bond length of 2.422 Å in **2** is not unusual for dimers of this type. The neopentyl groups and the chlorine atoms adopt a *trans* conformation about the gallium centers; and the coordination geometry about the Ga and P atoms approximates that of a distorted tetrahedron, with the mean Ga-P-Si and P-Ga-Cl angles being 113.1° and 104.9°, respectively. Dimer **3** crystallizes in the centrosymmetric space group $\text{P}2_1/\text{c}$ with the two molecules in the unit cell lying in general

positions. As in **2**, the alkyl groups and the bromine atoms adopt a *trans* conformation about the gallium centers. The mean Ga-P bond length of 2.424 Å in **3** is almost identical to that of 2.422 Å observed in **2**. The $\overline{\text{Ga-P-Ga'-P'}}$ ring is planar, and the coordination geometry about the Ga and P centers approximates a distorted tetrahedron [mean Ga-P-Si and P-Ga-Cl angles being 112.9° and 106.3°, respectively].

The average Ga-P bond length of 2.517 Å in **4** is slightly longer than those found in dimers **2**, **3**, and **13**. This may be attributed to the added steric bulk of the alkyl substituents on the gallium center. The Ga-P-Ga angle of 93.47(8)° and the P-Ga-P angle of 86.29(5)° evidence the degree of distortion from a perfect square for the four-membered ring. A comprehensive survey of the structural features observed in compound **4** revealed no unusual bond lengths or angles.

Compounds **5** and **6** crystallize in the monoclinic space group $P2_1/n$, with four molecules per unit cell. The asymmetric unit in **5** contains one molecule occupying a general position in the cell, and the four-membered $\overline{\text{Ga-P-Ga-Cl}}$ ring is nearly planar. This is in contrast to the four-membered ring in **6** where the dihedral angle between the plane defined by Ga(1)-P(1)-Cl(1) and Ga(2)-P(1)-Cl(1) in **6** is 23°. This angle is notably larger than the 6° dihedral angle observed in (**7**) and may be attributed to steric effects. It is interesting to note that the degree of four-membered ring planarity observed in **5** and **6** matches closely with that observed in the As analogs of these compounds²¹. The As analog of **5** has a mean dihedral angle of 0.05°, while the analog of **6** has a mean dihedral angle of 21.1°. In **6**, the Ga(1)-P-Ga(2) angle at 90.9(1)° and the Ga(1)-Cl-Ga(2) angle at 92.5(1)° evidence a slight degree of distortion for the four-membered ring. The exocyclic Si-P-Si and average C-Ga-C angles of 105.8(1)° and 125.75°, respectively, show the distorted tetrahedral coordination geometry about the phosphorus and gallium atoms. There were no uncharacteristically short intermolecular contacts or other unusual structural features observed in the crystal packing.

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Supplementary Material Available: Tables of bond distances and bond angles, anisotropic temperature factor parameters, hydrogen atom fractional coordinates, and structure amplitudes for **1 - 6** (143 pages). Ordering information is given on any current masthead page.

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Captions to Figures

Figure 1. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of **1**. Hydrogen atoms are omitted for clarity.

Figure 2. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of **2**. Hydrogen atoms are omitted for clarity. Primed atoms are related to non-primed atoms by a center of symmetry.

Figure 3. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of **3**. Hydrogen atoms are omitted for clarity. Primed atoms are related to non-primed atoms by a center of symmetry.

Figure 4. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of **4**. Hydrogen atoms are omitted for clarity.

Figure 5. ORTEP diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of **5**. Hydrogen atoms are omitted for clarity.

Figure 6. ORTEP diagram (35% probability ellipsoids) showing the solid-state structure and atom numbering scheme of **6**. Hydrogen atoms are omitted for clarity.

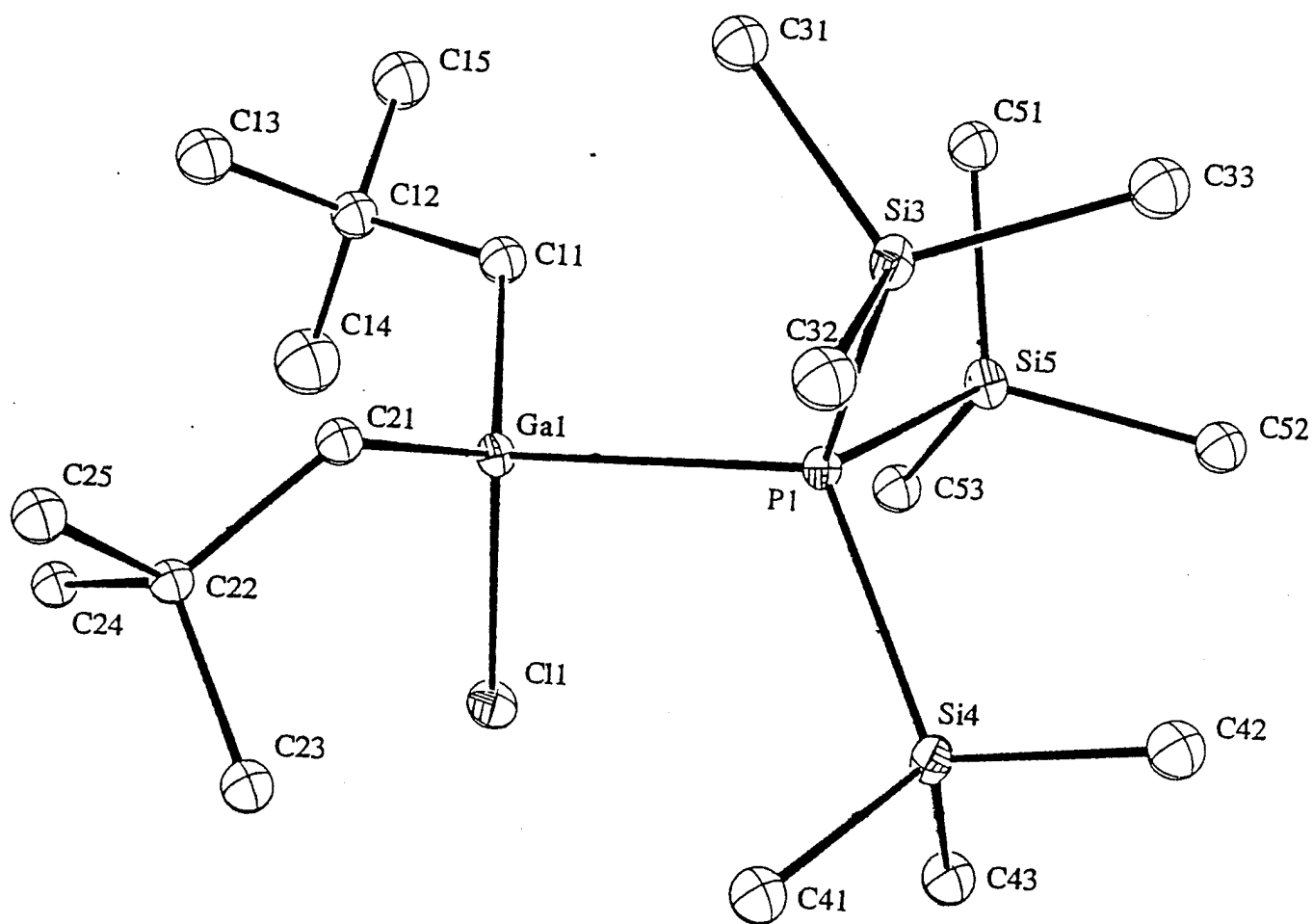


Figure 1

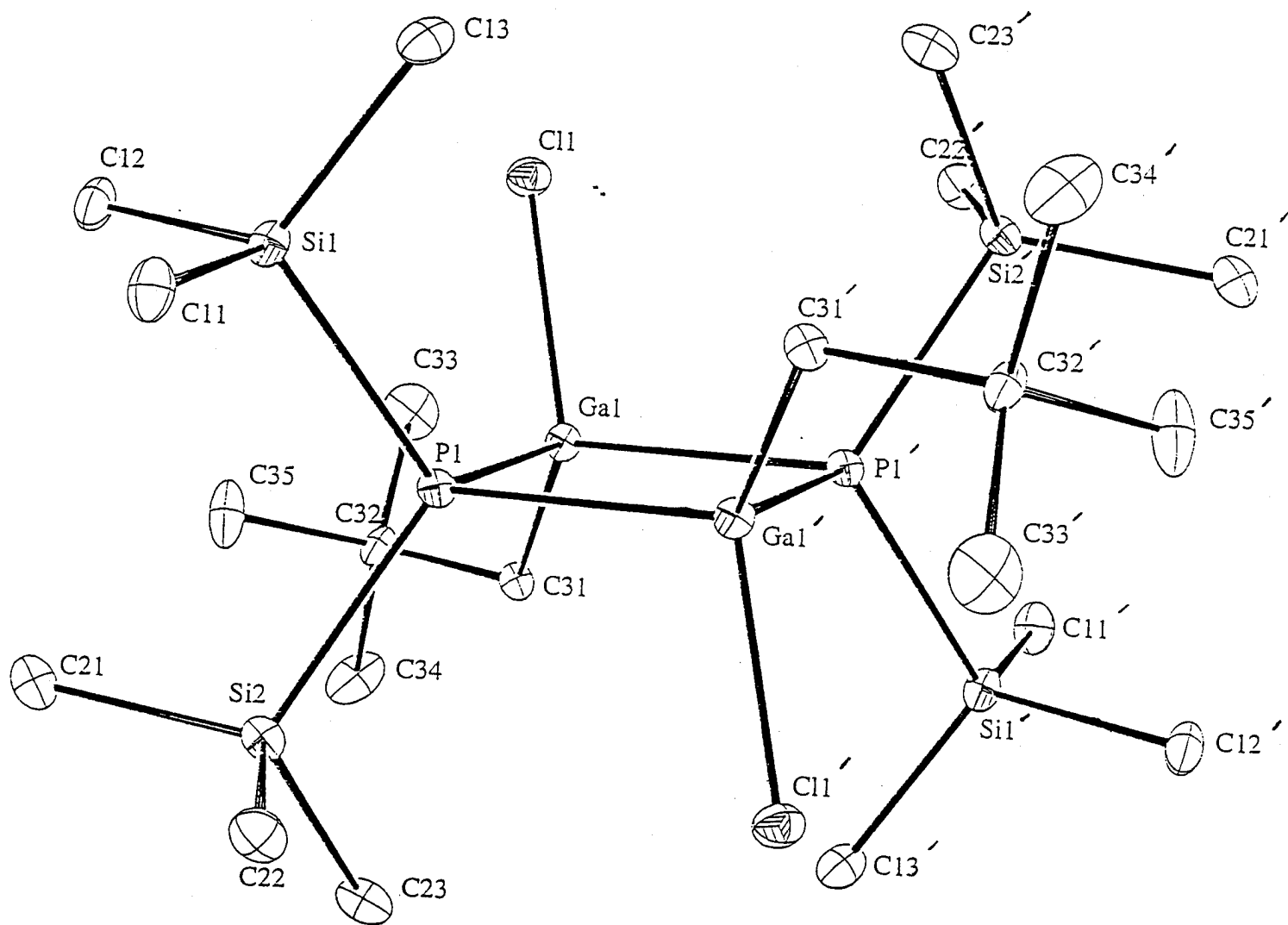


Figure 2

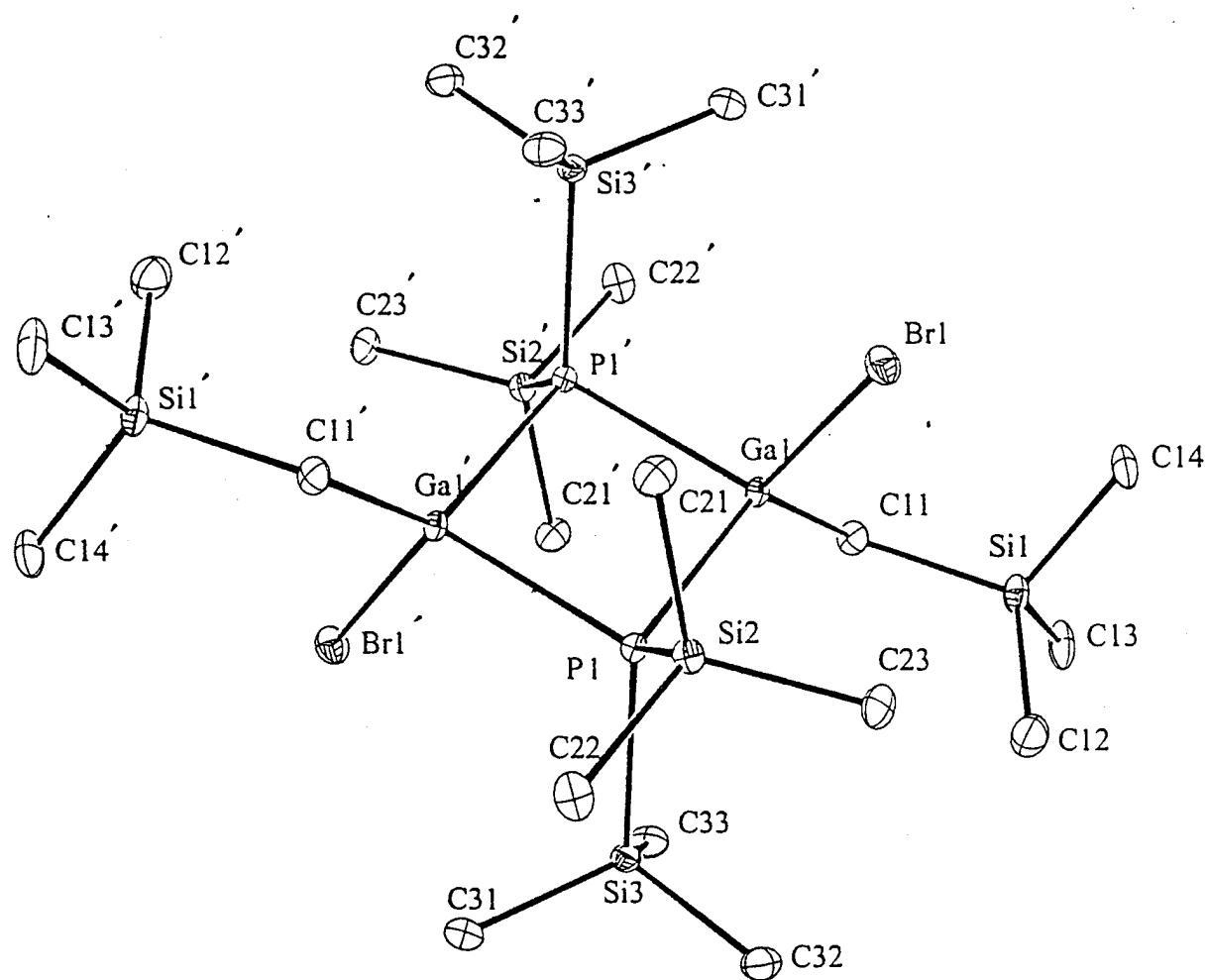


Figure 3

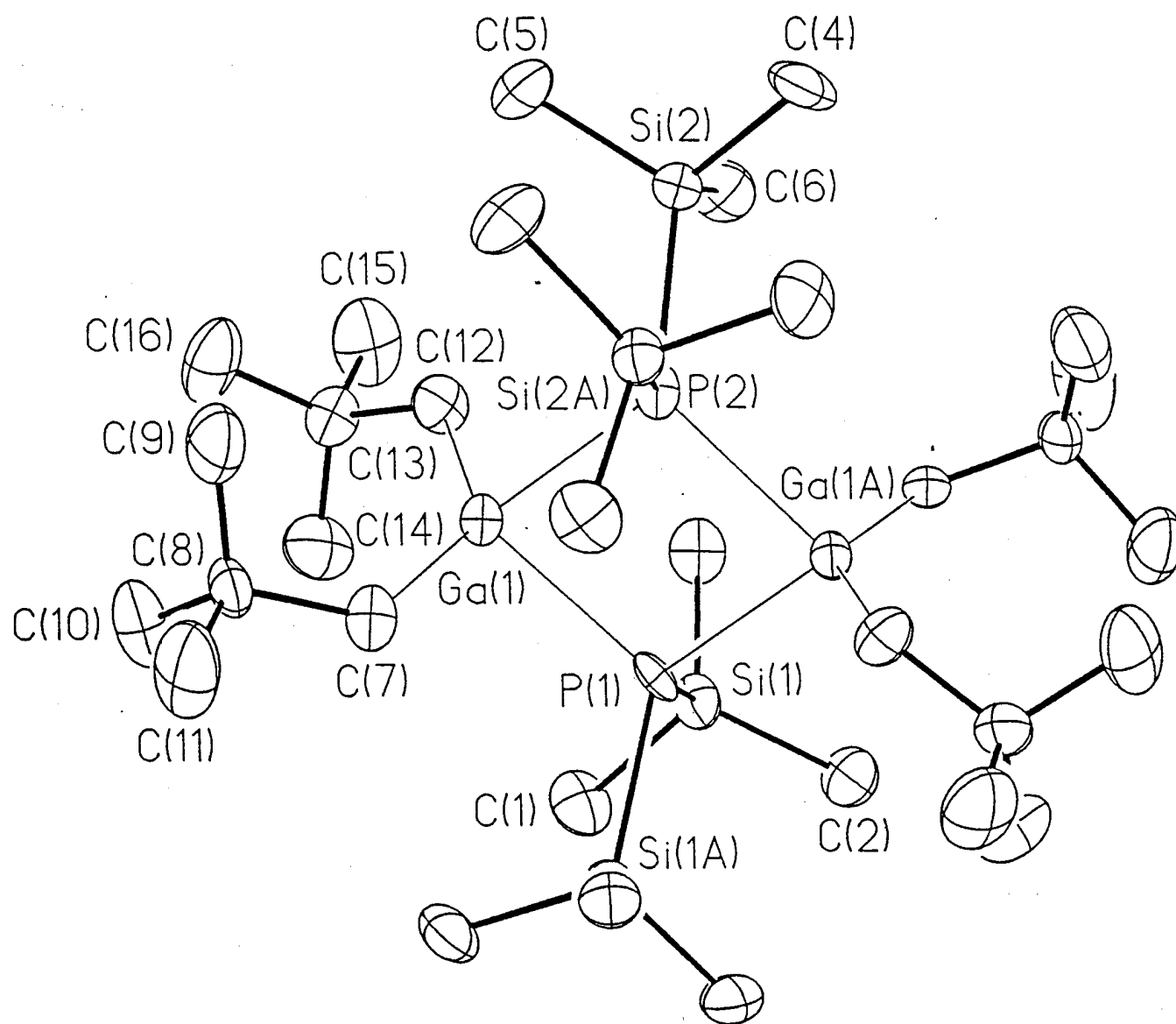


Figure 4

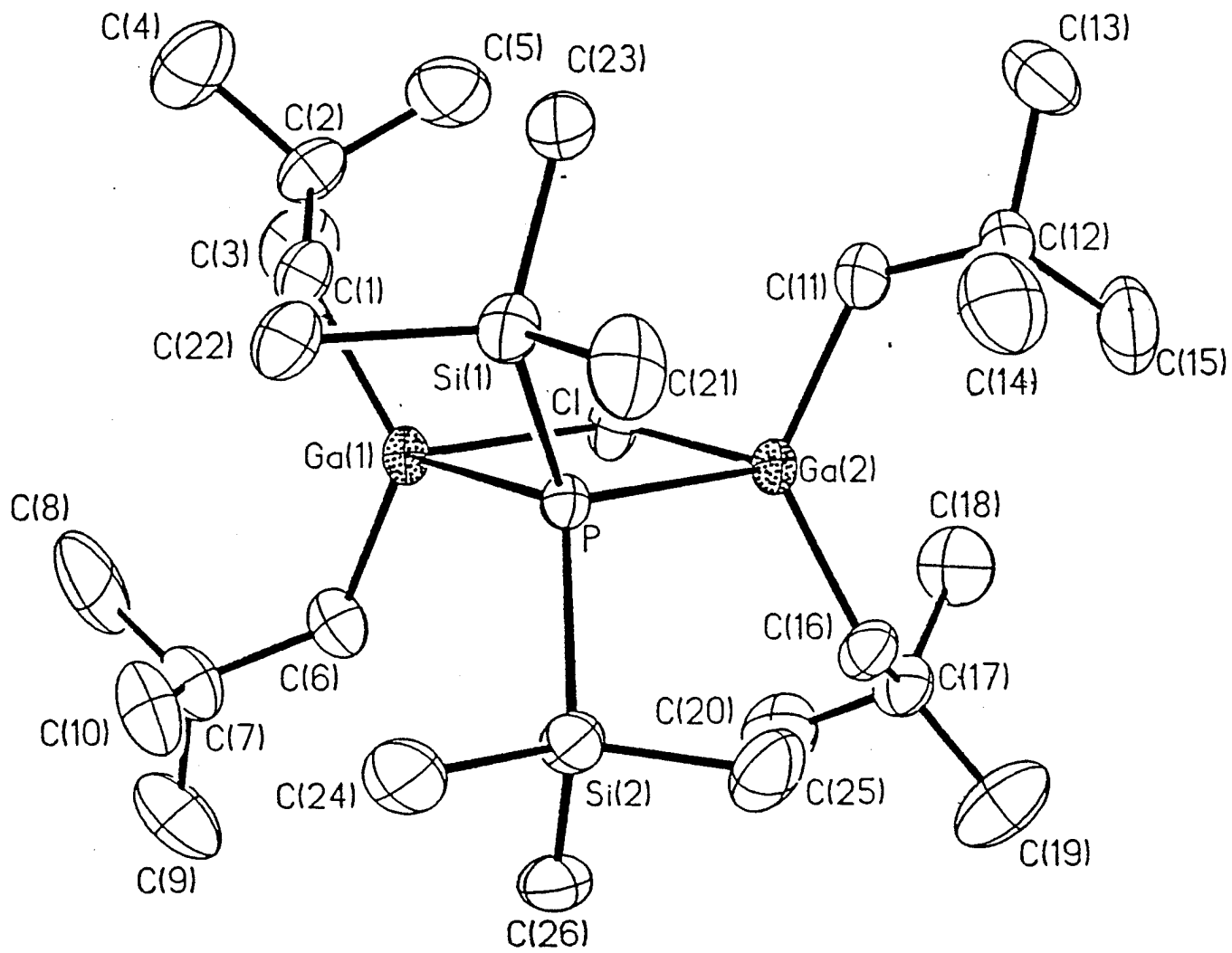


Figure 5

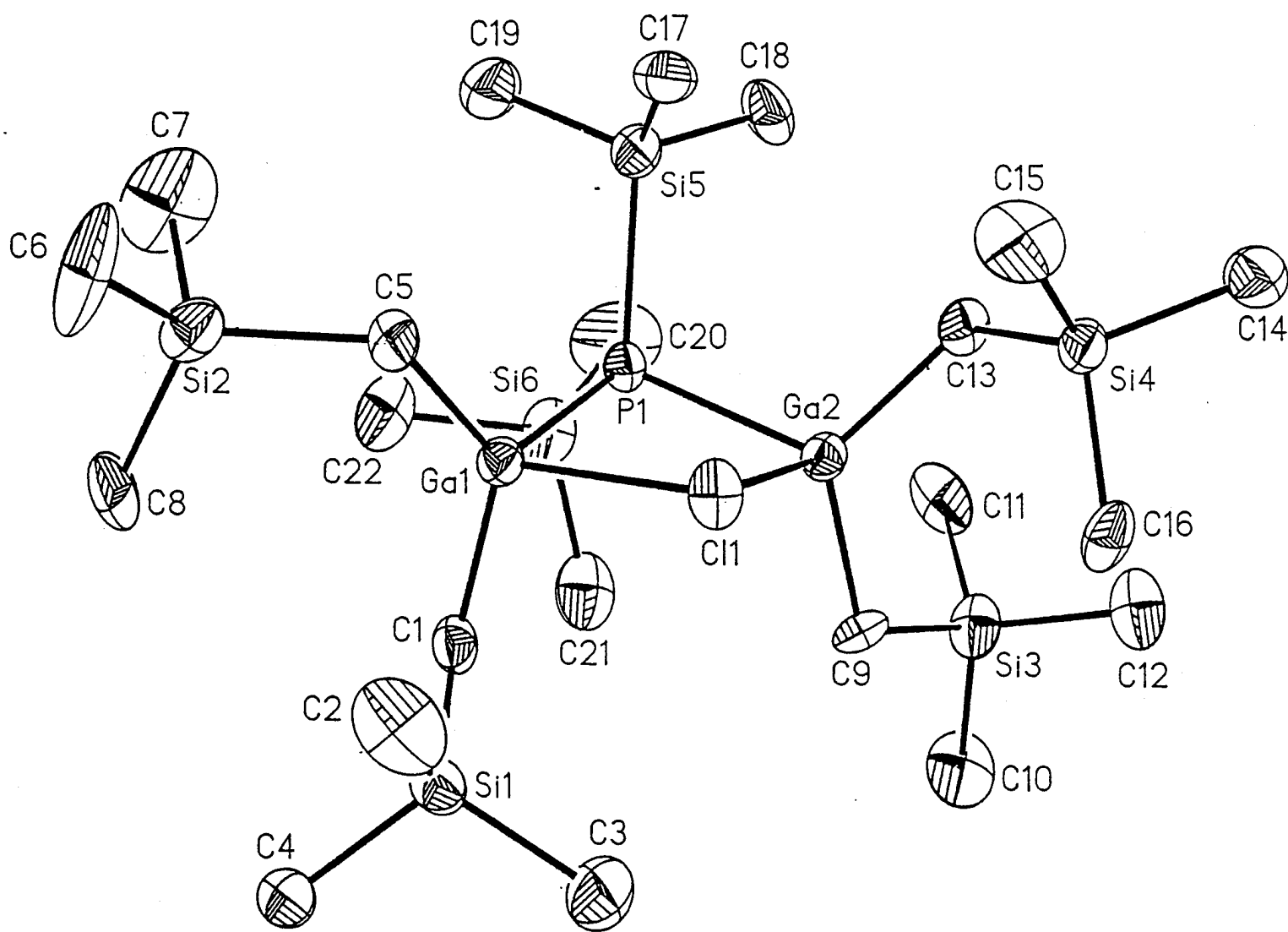


Figure 6

Table 1. Crystallographic Data and Measurements for $(\text{Me}_3\text{CCH}_2)_2\text{Ga}(\text{Cl})\cdot\text{P}(\text{SiMe}_3)_3$ (1), $[(\text{Me}_3\text{CCH}_2)(\text{Cl})\text{GaP}(\text{SiMe}_3)_2]_2$ (2), $[(\text{Me}_3\text{SiCH}_2)(\text{Br})\text{GaP}(\text{SiMe}_3)_2]_2$ (3), $[(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2]_2$ (4), $(\text{Me}_3\text{CCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{CMe}_3)_2\text{Cl}$ (5), and $(\text{Me}_3\text{SiCH}_2)_2\text{GaP}(\text{SiMe}_3)_2\text{Ga}(\text{CH}_2\text{SiMe}_3)_2\text{Cl}$ (6)

	1	2	3
molecular formula	$\text{C}_{19}\text{H}_{49}\text{ClGaPSi}_3$	$\text{C}_{22}\text{H}_{58}\text{Cl}_2\text{Ga}_2\text{P}_2\text{Si}_4$	$\text{C}_{20}\text{H}_{58}\text{Br}_2\text{Ga}_2\text{P}_2\text{Si}_6$
formula weight	497.99	707.32	828.38
crystal system	monoclinic	monoclinic	monoclinic
space group	$\text{P } 2_1/\text{c}$	$\text{P } 2_1/\text{c}$	$\text{P } 2_1/\text{c}$
a, Å	12.320(7)	12.000(4)	9.705(4)
b, Å	12.308(5)	10.191(5)	19.020(5)
c, Å	19.550(8)	17.402(3)	11.268(4)
β , deg	101.97(4)	92.910(20)	111.61(3)
V, Å ³	2900.0(23)	2125.4(13)	1933.8(11)
Z	4	2	2
radiation	Mo-K α (0.71073)	Mo-K α (0.71073)	Mo-K α (0.71073)
(wavelength, Å)			
μ , cm ⁻¹	12.2	16.6	37.1
temp, °C	-135	-135	-135
D _{calcd} , g cm ⁻³	1.141	1.273	1.423
crystal	0.40 x 0.35 x 0.35	0.30 x 0.30 x 0.20	0.30 x 0.30 x 0.25
dimens., mm			
T _{max} ; T _{min}	0.725:0.577	0.707:0.469	0.450:0.289
scan type	ω	ω	ω
Θ_{max} , deg	45	50	50
no. of rflns recorded	3768	3818	3782

Table 1 (continued)

	1	2	3
no. of non-equiv.	3762	3730	3408
rflns recorded			
R_{merg} (on I)	0.057	0.058	0.025
no. of rflns retained, $I > 2.5\sigma(I)$ or $I > 3.0\sigma(I)$	2827	2380	2387
no. of params. refined	132	174	145
R ; R_w ^a	0.043; 0.051	0.063; 0.075	0.036; 0.041
goodness-of-fit ^b	1.55	2.06	1.19
max shift / esd. in final least-squares cycle	0.000	0.030	0.004
final max, min $\Delta\rho$, e/Å ⁻³	0.740; -0.960	1.780; -1.730	0.620; -0.740

$$^a R = \Sigma(|F_o| - |F_c|) / \Sigma |F_o| ; R_w = [\Sigma w (|F_o| - |F_c|)^2 / \Sigma w |F_o|^2]^{1/2}.$$

$$^b \text{Goodness-of-fit} = [\Sigma w \Delta^2 / (N_{\text{observations}} - N_{\text{parameters}})]^{1/2}.$$

Table 1 (continued)

	4	5	6
molecular formula	C ₃₂ H ₈₀ Ga ₂ P ₂ Si ₄	C ₂₆ H ₆₂ ClGa ₂ PSi ₂	C ₂₂ H ₆₂ ClGa ₂ PSi ₆
formula weight	778.7	636.8	701.12
crystal system	monoclinic	monoclinic	monoclinic
space group	P 2/n	P 2 ₁ /n	P 2 ₁ /n
a, Å	12.1069(7)	12.355(1)	16.149(5)
b, Å	12.5860(13)	17.257(2)	23.144(5)
c, Å	15.321(2)	17.713(3)	10.700(3)
β, deg	90.148(8)	98.02(1)	90.88(2)
V, Å ³	2334.6(4)	3739.7(6)	3999(1)
Z	2	4	4
radiation	Mo-Kα (0.71073)	Mo-Kα (0.71073)	Mo-Kα (0.71073)
(wavelength, Å)			
μ, cm ⁻¹	13.43	16.31	16.4
temp, °C	25	25	-140
D _{calcd} , g cm ⁻³	1.108	1.131	1.16
crystal	0.40 x 0.40 x 0.30	0.30 x 0.30 x 0.40	0.18 x 0.34 x 0.48
dimens., mm			
T _{max} ; T _{min}	---	---	0.69:1.00
scan type	ω	ω	ω/2θ
Θ _{max} , deg	60	45	50
no. of rflns recorded	6707	7631	8634
no. of non-equiv.	5211	4844	7024
rflns recorded			
R _{merg} (on I)	0.071	0.024	0.072

Table 1 (continued)

	4	5	6
no. of rflns retained, $I > 2.5\sigma(I)$ or $I > 3.0\sigma(I)$	2126	2654	3589
no. of params. refined	182	289	290
$R; R_w^a$	0.068; 0.177 ^c	0.038; 0.042	0.0805; 0.0806
goodness-of-fit ^b	0.960	1.00	3.42
max shift / esd. in final least-squares cycle	0.000	0.000	0.0006
final max, min $\Delta\rho$, e/Å ⁻³	0.551; -0.539	0.36; -0.30	0.89; -0.82

$$^aR = \Sigma(|F_o| - |F_c|)/\Sigma|F_o|; R_w = [\Sigma w(|F_o| - |F_c|)^2/\Sigma w|F_o|^2]^{1/2}.$$

$$^b\text{Goodness-of-fit} = [\Sigma w\Delta^2/(N_{\text{observations}} - N_{\text{parameters}})]^{1/2}.$$

$$^cR(wF^2) = [\Sigma[w(F_o^2 - F_c^2)^2/\Sigma(wF_o^2)]^{1/2}$$

Table 2. Selected bond distances (\AA) for **1** - **6**, with Estimated Standard Deviations in Parentheses

Bond Lengths			
1			
Ga-P(1)	2.534(2)	Ga-Cl(1)	2.258(2)
P(1)-Si(3)	2.275(2)	Ga-C(11)	1.998(5)
P(1)-Si(4)	2.279(2)	Ga-C(21)	1.999(5)
P(1)-Si(5)	2.285(2)		
2			
Ga-P(1)	2.419(2)	Ga-Cl(1)	2.240(2)
Ga-P(1)'	2.424(2)	Ga-C(31)	1.987(8)
P(1)-Si(1)	2.275(3)		
P(1)-Si(2)	2.281(3)		
3			
Ga-P(1)	2.427(2)	Ga-Br(1)	2.372(9)
Ga-P(1)'	2.421(2)	Ga-C(11)	1.979(5)
P(1)-Si(2)	2.272(2)		
P(1)-Si(3)	2.278(2)		
4			
Ga-P(1)	2.512(2)	Ga-C(12)	2.000(7)
Ga-P(2)	2.522(2)		
P(1)-Si(1)	2.267(2)		
P(2)-Si(2)	2.264(2)		

Table 2. (continued)

5

Ga(1)-Cl(1)	2.420 (2)	Ga(1)-P(1)	2.446 (2)
Ga(1)-C(1)	1.966 (6)	Ga(2)-P(1)	2.456 (2)
Ga(2)-Cl(1)	2.415 (2)	Ga(2)-C(16)	1.976 (6)
Ga(2)-C(11)	1.959 (6)		

6

Ga(1)-Cl(1)	2.435 (4)	Ga(1)-P(1)	2.414 (4)
Ga(1)-C(1)	1.97 (1)	Ga(2)-P(1)	2.418 (4)
Ga(2)-Cl(1)	2.419 (4)	Ga(2)-C(13)	1.97 (1)
Ga(2)-C(9)	1.98 (1)		

Table 3. Selected bond angles (°) for **1** - **6**, with Estimated Standard Deviations in Parentheses

Bond Angles			
1			
Ga - P(1) - Si(3)	110.34(7)	P(1) - Ga - Cl(1)	97.49(6)
Ga - P(1) - Si(4)	114.87(7)	P(1) - Ga - C(11)	101.51(16)
Ga - P(1) - Si(5)	111.80(7)	P(1) - Ga - C(21)	106.23(15)
Si(3) - P(1) - Si(4)	107.14(8)	Cl(1) - Ga - C(11)	110.84(15)
Si(3) - P(1) - Si(5)	107.10(7)	Cl(1) - Ga - C(21)	115.63(15)
Si(4) - P(1) - Si(5)	105.15(7)	C(11) - Ga - C(21)	120.99(21)
2			
Ga(1) - P(1) - Ga(1)'	89.92(8)	P(1)' - Ga(1) - Cl(1)	104.76(8)
Ga(1) - P(1) - Si(1)	113.09(10)	P(1)' - Ga(1) - C(31)	121.39(24)
Ga(1) - P(1) - Si(2)	113.09(11)	Ga(1)' - P(1) - Si(1)	113.60(11)
Si(1) - P(1) - Si(2)	108.31(12)	Ga(1)' - P(1) - Si(2)	118.01(10)
P(1) - Ga(1) - P(1)'	90.08(7)		
P(1) - Ga(1) - Cl(1)	105.01(8)		
3			
Ga(1) - P(1) - Ga(1)'	90.32(5)	P(1)' - Ga(1) - Br(1)	104.52(4)
Ga(1) - P(1) - Si(2)	113.96(7)	P(1)' - Ga(1) - C(11)	124.52(16)
Ga(1) - P(1) - Si(3)	112.01(7)	Ga(1)' - P(1) - Si(2)	113.80(7)
Si(2) - P(1) - Si(3)	107.36(8)	Ga(1)' - P(1) - Si(3)	118.85(7)
P(1) - Ga(1) - P(1)'	89.68(5)		
P(1) - Ga(1) - Br(1)	108.14(4)		

Table 3. (continued)

4

Ga(1) - P(2) - Ga(1A)	93.47(8)	C(7) - Ga(1) - P(1)	103.8(2)
Ga(1) - P(1) - Si(1)	118.41(5)	Si(1) - P(1) - Ga(1)	118.41(5)
Si(2) - P(2) - Si(2A)	103.76(14)		
P(1) - Ga(1) - P(2)	86.29(5)		
C(7) - Ga(1) - C(12)	133.0(3)		

5

Cl(1)-Ga(1)-P(1)	88.4(1)	Cl(1)-Ga(1)-C(1)	109.0(2)
P(1)-Ga(1)-C(1)	112.5(2)	Cl(1)-Ga(1)-C(6)	98.2(2)
P(1)-Ga(1)-C(6)	114.9(2)	Cl(1)-Ga(2)-C(11)	98.1(2)
Cl(1)-Ga(2)-P(1)	88.3(1)	Cl(1)-Ga(2)-C(16)	108.7(2)
P(1)-Ga(2)-C(11)	115.4(2)	Ga(1)-P(1)-Ga(2)	90.9(1)
Ga(1)-Cl(1)-Ga(2)	92.5(1)	Ga(2)-P(1)-Si(1)	120.6(1)
Ga(1)-P(1)-Si(1)	108.9(1)	Ga(2)-P(1)-Si(2)	110.3(1)
Ga(1)-P(1)-Si(2)	120.8(1)		

6

Cl(1)-Ga(1)-P(1)	87.3(1)	Cl(1)-Ga(1)-C(1)	102.7(4)
P(1)-Ga(1)-C(1)	108.7(4)	Cl(1)-Ga(1)-C(5)	102.2(4)
P(1)-Ga(1)-C(5)	119.5(4)	C(1)-Ga(1)-C(5)	126.1(6)
Cl(1)-Ga(2)-P(1)	87.6(1)	Cl(1)-Ga(2)-C(9)	105.1(4)
P(1)-Ga(2)-C(9)	114.2(4)	Cl(1)-Ga(2)-C(13)	104.7(4)
Ga(1)-Cl(1)-Ga(2)	89.9(1)	Ga(1)-P(1)-Ga(2)	90.4(1)
Ga(1)-P(1)-Si(5)	116.0(2)	Ga(2)-P(1)-Si(5)	109.9(2)
Ga(1)-P(1)-Si(6)	110.2(2)	Ga(2)-P(1)-Si(6)	121.3(2)

Table 4. Selected Non-Hydrogen Atom Fractional Coordinates and Equivalent Isotropic Thermal Parameters for 1 - 3, with Estimated Standard Deviations in Parentheses

Atom	x	y	z	B _{iso} (Å ²) ^a
1				
Ga1	0.71428(4)	0.39355(4)	0.13862(3)	1.19(2)
Cl1	0.83862(11)	0.41926(11)	0.23972(6)	1.91(6)
P1	0.81451(11)	0.24006(10)	0.09134(6)	1.12(5)
Si3	0.75421(12)	0.22521(11)	-0.02645(7)	1.43(6)
Si4	1.00269(12)	0.25718(11)	0.11215(7)	1.44(6)
Si5	0.78298(13)	0.07649(11)	0.13895(7)	1.53(6)
C11	0.5791(4)	0.3177(4)	0.1559(3)	1.81(10)
C12	0.4938(4)	0.3760(4)	0.1898(3)	1.99(10)
C13	0.4444(5)	0.4722(5)	0.1454(3)	2.65(11)
C14	0.5488(6)	0.4160(6)	0.2627(4)	4.05(15)
C15	0.4004(5)	0.2968(5)	0.1958(3)	3.19(13)
C21	0.7022(4)	0.5143(4)	0.06906(25)	1.53(9)
C22	0.7427(4)	0.6312(4)	0.08905(25)	1.37(9)
C23	0.8692(4)	0.6345(4)	0.1089(3)	1.77(10)
2				
Ga1	0.89428(7)	0.93924(9)	0.05199(5)	1.01(3)
P1	1.09586(17)	0.93294(21)	0.06496(11)	1.02(8)
Cl1	0.84365(17)	0.73805(20)	0.01196(12)	1.64(8)
Si1	0.5346(8)	0.3894(9)	0.1330(5)	6.0(4)
Si2	1.16632(19)	1.03857(23)	0.17187(12)	1.28(9)
C11	1.16599(20)	0.72490(22)	0.06558(12)	1.23(9)
C12	1.1501(7)	0.9296(9)	0.2572(5)	1.9(4)
C13	1.3164(7)	1.0746(9)	0.1596(5)	1.9(4)
3				
Ga1	0.61904(6)	0.46319(3)	0.13939(5)	1.075(23)
Br1	0.49747(6)	0.39270(3)	0.24465(5)	1.966(24)
P1	0.48623(14)	0.57379(7)	0.08155(12)	1.03(5)
Si1	0.94569(15)	0.39569(8)	0.32154(12)	1.67(6)
Si2	0.63588(15)	0.66950(8)	0.11463(13)	1.29(6)
Si3	0.33204(15)	0.59174(8)	0.18980(13)	1.25(6)

Table 4 continued.

Atom	x	y	z	$B_{\text{iso}}(\text{\AA}^2)^a$
3				
C11	0.8338(5)	0.4733(3)	0.2386(5)	1.61(22)
C12	0.9213(7)	0.3235(3)	0.2029(6)	2.6(3)
C13	1.1466(6)	0.4185(3)	0.3873(6)	2.6(3)
C14	0.8941(6)	0.3658(3)	0.4573(5)	2.2(3)
C21	0.7975(6)	0.6520(3)	0.0670(5)	2.0(3)
C22	0.7079(6)	0.6909(3)	0.2881(5)	1.98(25)

^a B_{iso} = the mean of the principal axes of the thermal ellipsoid

Table 5. Selected Non-Hydrogen Atom Fractional Coordinates ($\times 10^4$) and Equivalent Isotropic Displacement Coefficients ($\text{\AA}^2 \times 10^3$) for **4** - **6**, with Estimated Standard Deviations in Parentheses

Atom	x	y	z	U_{eq}^*
4				
Ga(1)	8232(1)	7458(1)	1451(1)	41(1)
P(1)	7500	8819(2)	2500	38(1)
P(2)	7500	6084(2)	2500	38(1)
Si(1)	6153(2)	9937(2)	2033(1)	52(1)
Si(2)	6252(2)	4974(2)	1882(1)	48(1)
C(1)	6723(7)	11036(6)	1342(5)	81(3)
C(2)	5428(7)	10589(6)	2987(5)	73(2)
C(3)	5098(7)	9224(6)	1384(5)	71(2)
C(4)	5817(8)	3920(6)	2674(5)	84(3)
C(5)	6832(7)	4272(6)	905(4)	73(2)
5				
Ga(1)	-37.9(6)	1265.9(4)	7058.9(4)	46(1)
Ga(2)	1479.8(6)	2431.0(4)	5982.1(4)	49(1)
Cl(1)	1067(2)	1067(1)	6049(1)	67(1)
P(1)	350(1)	2653(1)	6997(1)	40(1)
Si(1)	-1253(2)	3308(1)	6753(1)	56(1)
Si(2)	1334(2)	3271(1)	7997(1)	57(1)
C(1)	-1568(5)	1018(4)	6673(4)	60(3)
C(2)	-1890(6)	392(4)	6061(5)	69(3)
C(3)	-1214(7)	-331(4)	6225(5)	101(4)
C(4)	-3091(7)	180(5)	6057(5)	121(5)
C(5)	-1739(7)	696(5)	5276(4)	102(4)
6				
Ga(1)	2998(1)	140(1)	6418(2)	26(1)
Ga(2)	2245(1)	1455(1)	5458(1)	26(1)
Cl(1)	3405(2)	855(2)	4877(3)	33(1)
P(1)	2150(2)	881(2)	7338(3)	25(1)
Si(1)	2720(3)	-719(2)	3996(4)	37(2)
Si(2)	4123(3)	-772(2)	8144(4)	39(2)
Si(3)	522(3)	1859(2)	4129(5)	41(2)
Si(4)	3448(3)	2507(2)	4520(4)	31(1)
Si(5)	2721(3)	1362(2)	8964(4)	35(2)
Si(6)	928(3)	501(2)	7951(4)	39(2)
C(1)	2281(8)	-382(6)	5437(13)	32(5)

Table 5 continued.

Atom	x	y	z	U_{eq}^*
6				
C(2)	3866(9)	-741(7)	3985(16)	67(8)
C(3)	2329(10)	-309(7)	2588(14)	54(7)
C(4)	2389(10)	-1495(6)	3889(14)	55(7)
C(5)	4057(7)	-61(6)	7206(14)	32(5)

* Equivalent isotropic U defined as one third of the trace of the orthogonalized U_{ij} tensor

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